

CODE NO: 00000

**T.I. MATRICULATION HIGHER SECONDARY SCHOOL, AMBATTUR
SUMMATIVE ASSESSMENT II- 2018**

MARKS: 30

ROLL NO :

TIME :

I Choose the correct answer :- (Each 1 mark) 4 x 1=4

1. Which of the following will gain electrons in a chemical reaction?
a) F
2. Which among these is not an oxidation reaction?
a) $2\text{Na} + \text{H}_2 \rightarrow 2\text{NaH}$
3. Which among these are coordinate covalent compounds?
a) benzene
4. The maximum number of electrons that can be accommodated in s,p,d,f subshells are :-
b) 2,6,10,14

II Give two examples for each:- (Any two-each ½ marks) 2x1=2

- 1) Alkali metals-
- 2) Halogens

III Answer any seven of the following : 7 x 2=14

1. State law of Dobereiner's law of triads.
The atomic mass of the middle element is nearly the same as the average of the atomic masses of the other two elements.
2. Why were the elements of group 18 known as inert gases?
These elements do not react with other substances due to their completely filled shells. so they are called inert gases.
3. Look at the example and complete the table: Each answer-1/2 mark

S.No.	Alloy	Composition	One use
Eg.	Gun metal	Copper, tin and Zinc	Frame of spectacles
1	Steel	Iron and carbon	Ship construction
2	Bronze	Cu and Sn	Statue or medal

4. Which of the following statements are true?
c) Metals get reduced on reaction with non metals.
5. Draw the Lewis dot structure for the following elements:- Each diagram ½ marks
a) Nitrogen b) Aluminium c) Hydrogen d) Sodium

6. Complete the following table:- Each ½ marks

S.No.	Element	Electron distribution	Valence electrons
1.	Carbon	2,4	4
2.	Calcium	2,8,8,2	2

7. Give reasons for the following:- each-1 mark

a) Covalent compounds are bad conductors of electricity.

They do not contain charged particles.

b) Ionic compounds have very high melting and boiling points.

Strong electrostatic force between cations and anions hold them tightly together.

8. How are ionic compounds different from covalent compounds? List any four points to support your answer.any 4,each point-1/2 marks

IV. Answer any two of the following:

(2x5=10)

1. Draw the diagram to represent the bonding in Magnesium chloride. Briefly explain the type of bond.

Diagram of Mg-1 mark

Cl-1 mark

,MgCl₂molecule-1 mark

Explanation of ionic bond by transfer of 2electrons from Mg to each Cl-2 marks

2. What is meant by coordinate covalent bond? Explain with an example.

Explanation of the bond-2 marks

Example explanation-1 mark

Diagram-2 marks

3. Write a short note on each of the following:- (2 ½ marks each-1 mark for reason for calling them as s block,1 mark for two properties, ½ for two examples)

a) s block elements-

b) d block elements

4. Write the features of the modern periodic table.

Any five points-each 1 mark
